

## Read PDF Chapter 9 Review Stoichiometry Section 2 Work

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### **Chapter 9 Review Stoichiometry Section**

CHAPTER 9 REVIEW Stoichiometry  
SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.
2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the following equation:  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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CHAPTER 9 REVIEW Stoichiometry  
SECTION 3 PROBLEMS Write the answer  
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mixed with 12.0 mol of H

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## **Chapter 9: Stoichiometry Review and Chapter Summary ...**

Stoichiometry. SECTION 1. SHORT  
ANSWER Answer the following questions  
in the space provided. 1. \_\_\_\_\_ The  
coefficients in a chemical equation  
represent the (a) masses in grams of all

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reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the following

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CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of  $N_2$  are mixed with 120 mol of H

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## **[PDF] Chapter 9 Stoichiometry Section 2 Worksheet**

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SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:  $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How many grams of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed? 2. Given the following equation ...

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Reaction stoichiometry uses molar

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Review Module / Chapters 9-12 13 Prentice Hall, Inc. All rights In your notebook,solve the following problems. SECTION 9.1 THE ARITHMETIC OF EQUATIONS Use the 3-step problem-solving approach you learned in Chapter 4. 1. An apple pie needs 10 large apples, 2 crusts (top and bottom), and 1 tablespoon of cinnamon.

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## 9 Stoichiometry Practice Problems

CHAPTER 9 REVIEW Stoichiometry  
SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided

1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield

2 60 mol of  $N_2$  are mixed with 120 mol of  $H_2$  according to the following equation:  $N_2(g) + 3H_2(g)$

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Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

## Stoichiometry Worksheet Answers

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## Chapter 9

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SECTION 2 PROBLEMS Write the answer  
on the line to the left. Show all your work  
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following equation represents a  
laboratory preparation for oxygen gas:  
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are totally consumed? ...

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