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3 Stoichiometry Solutions to Black Exercises
Total mass C is 0.200 g + 0.200 g = 0.400 g C. $O_{45}H_{20}$
 $1 \text{ mol } H_2O \times 2 \text{ mol } H \times 1.008 \text{ g } H : \text{---} =$
0.0503 g H 18.02 g H_2O 1 mol H_2O 1 mol

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H (Since hydrocarbons contain only the elements C and H, g H can also be obtained by subtraction: 0.450 g sample — 0.400 g C = 0.050 g H.) 0.400 g mi = 0.0333 mol C; $0.0333 / 0.0333 = 1.0$
12.01 g C 0.0503 g Hx 1i1H— = 0.0499 mol H; $0.0499 / 0.0333 = 1.5$ 1.008 g H
Multiplying by a factor of 2, the empirical ...

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11th 1 Matter and Measurement
Solutions to Black Exercises 1.71 (a)
density = $(35.66 \text{ g} - 14.23 \text{ g})/4.59 \text{ cm}^3$
 $= 4.57 \text{ g/ cm}^3$ 1000g \times 1mL 1L b 34.51<
— U 3 X 1 kg 13.6g X 1000 mL = 2.54
L (c) $V = 4/3 \pi r^3 = 4/3 \pi (28.9 \text{ cm})^3$
 $= 1.0111 \times 10^5 = 1.01 \times 10^5 = 1.01 \times \dots$

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